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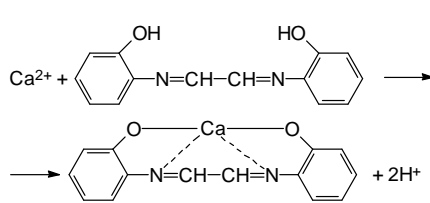
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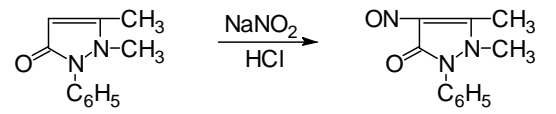
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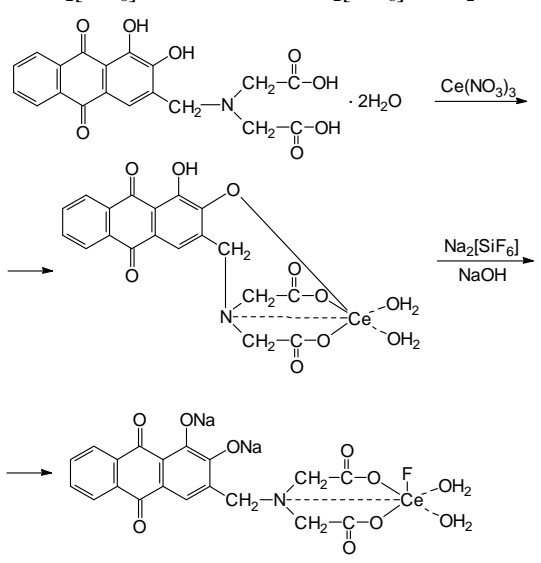
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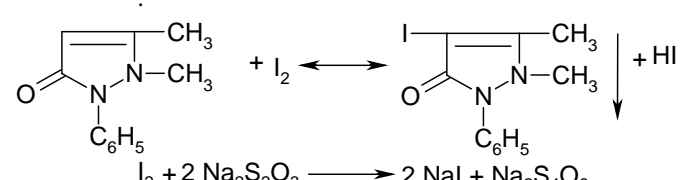
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| <p>1.</p> <p>;</p> <p>?</p> <p>A.</p> <p>B.</p> <p>D.</p> <p>E.</p> | <p>:</p> $\text{NH}_4^+ + \text{NaOH} \xrightarrow{t^\circ} \text{NH}_3\uparrow + \text{H}_2\text{O} + \text{Na}^+$ |
| <p>2.</p> <p>?</p> <p>A.</p> <p>(III)</p> <p>D.</p> <p>E.</p> | <p>C</p> $\text{SO}_3^{2-} + \text{I}_2 + \text{H}_2\text{O} \longrightarrow \text{SO}_4^{2-} + 2\text{HI}$ |
| <p>3.</p> <p>i - i</p> <p>i i</p> <p>i i</p> <p>A.</p> <p>(II)</p> <p>D.</p> <p>E.</p> | <p>-</p> $\text{Na}_2\text{S}_2\text{O}_3 + 2\text{AgNO}_3 \rightarrow \text{Ag}_2\text{S}_2\text{O}_3\downarrow + 2\text{NaNO}_3$ <p>()</p> $\text{Ag}_2\text{S}_2\text{O}_3 \rightarrow \text{Ag}_2\text{SO}_3\downarrow + \text{S}\downarrow$ <p>()</p> $\text{Ag}_2\text{SO}_3 + \text{S} + \text{H}_2\text{O} \rightarrow \text{Ag}_2\text{S}\downarrow + \text{H}_2\text{SO}_4$ <p>()</p> |
| <p>4.</p> <p>i i</p> <p>i i</p> <p>i i</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>(II)</p> <p>D.</p> <p>(III)</p> <p>E.</p> | <p>i i i i</p> <p>i</p> <p>:</p> $\text{Ca}^{2+} + 2\text{NH}_4\text{Cl} + \text{K}_4[\text{Fe}(\text{CN})_6] \rightarrow (\text{NH}_4)_2\text{Ca}[\text{Fe}(\text{CN})_6]\downarrow + 4\text{K}^+ + 2\text{Cl}^-$ |
| <p>5.</p> <p>i i i i</p> <p>i i i i</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | <p>:</p>  |
| <p>6.</p> <p>-i</p> <p>i</p> <p>i</p> <p>A.</p> <p>B.</p> <p>C.</p> | <p>-</p> <p>:</p> $\text{Cl}^- + \text{AgNO}_3 \rightarrow \text{AgCl}\downarrow + \text{NO}_3^-$ $\text{AgCl} + 2\text{NH}_4\text{OH} \rightarrow [\text{Ag}(\text{NH}_3)_2]\text{Cl} + 2\text{H}_2\text{O}$ |

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| D. E. i | |
| 7. i - i i i i A. i B. i C. i D. i E. i | - : $SO_4^{2-} + BaCl_2 \rightarrow BaSO_4\downarrow + 2Cl^-$ |
| 8. i , i i i i i i i i i i A. i B. i C. i D. i E. (III) | : $Mg^{2+} + Na_2HPO_4 + NH_4 \rightarrow NH_4MgPO_4\downarrow + 2Na^+ + H_2O$ |
| 9. ? A. i B. i C. i D. i E. i | - : : $4Na_2CO_3 + 4MgSO_4 + 4H_2O \rightarrow 3MgCO_3 \cdot Mg(OH)_2 \cdot 3H_2O\downarrow + 4Na_2SO_4 + CO_2\uparrow$ $2HCO_3^- + MgSO_4 \rightarrow Mg(HCO_3)_2 + SO_4^{2-}$ $Mg(HCO_3)_2 \rightarrow MgCO_3\downarrow + CO_2\uparrow + H_2O$ |
| 10. i i i i i i i i i i A. i i B. i C. i i i i D. i i E. i | ():  |
| 11. - : A. i B. i C. i D. - E. i | (III) (V) - : $NaH_2PO_2 + HCl \rightarrow H_3PO_2 + NaCl$ $As_2O_3 + 3H_3PO_2 \rightarrow 2As\downarrow + 3H_3PO_3$ $As_2O_5 + 5H_3PO_2 \rightarrow 2As\downarrow + 5H_3PO_3$ |

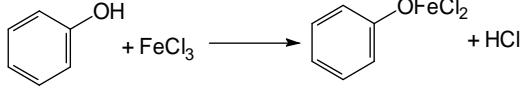
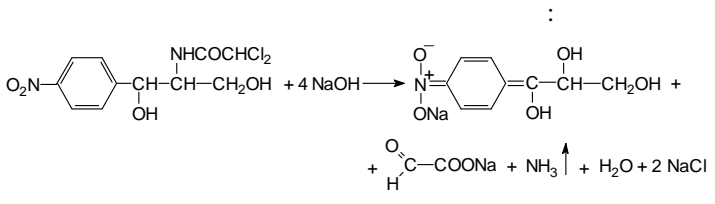
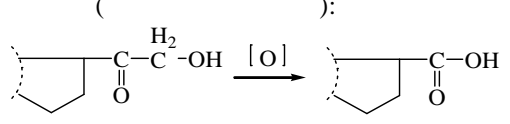
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| 1. Natrii iodidum Kalii iodidum : A. i | : $I_2 + 2Na_2S_2O_3 \rightarrow 2NaI + Na_2S_4O_6$ |

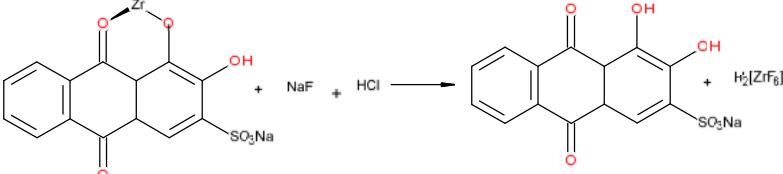
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| <p>D. E.</p> | |
| <p>2.</p> <p>A. D. E.</p> | <p style="text-align: right;">0,01</p> |
| <p>3.</p> <p>A. B. C. D. E.</p> | |
| <p>4.</p> <p>A. D. E.</p> | $\text{SO}_4^{2-} + \text{BaCl}_2 \xrightarrow{p.\text{CH}_3\text{COOH}} \downarrow \text{BaSO}_4 + 2 \text{Cl}^-$ |
| <p>5.</p> <p>A. B. C. D. 2,6- E.</p> | <p>(3,5):</p> $\text{CH}_3\text{-C} \begin{matrix} \text{S} \\ \parallel \\ \text{NH}_2 \end{matrix} + \text{H}_2\text{O} \longrightarrow \text{CH}_3\text{COO}^- + \text{NH}_4^+ + \text{H}_2\text{S} \uparrow$ $\text{H}_2\text{S} + \text{Pb}^{2+} \longrightarrow \text{PbS} \downarrow + 2\text{H}^+$ <p style="text-align: right;">(II)</p> |
| <p>6.</p> | <p>(</p> |

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| <p>i i i</p> <p>i :</p> <p>B. i i</p> <p>C. i i</p> <p>D. i</p> <p>E. i</p> | <p>SiO₂ + 6HF → H₂[SiF₆]↑ + 2H₂O</p> <p>H₂[SiF₆] + 2NaOH → Na₂[SiF₆] + 2H₂O</p>  |
| <p>7. i i - i</p> <p>i i i i</p> <p>i i :</p> <p>A. i i</p> <p>B.</p> <p>C. i</p> <p>D.</p> <p>E. i</p> | <p>Fe³⁺ + 2HS-CH₂COOH + 5NH₃·H₂O → [Fe(OH)(SCH₂COO⁻)₂]²⁻ + 5NH₄⁺ + 4H₂O</p> |

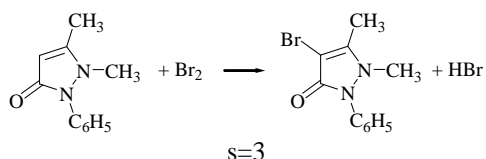
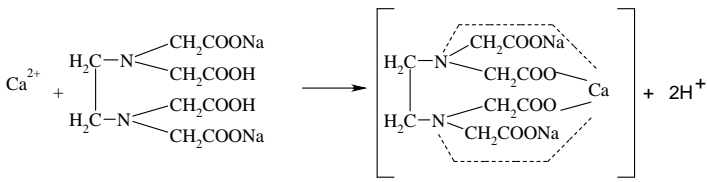
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| <p>« -2»</p> <p>1. i i - i</p> <p>i i i</p> <p>i</p> <p>i i ?</p> <p>A. i i</p> <p>B. i i</p> <p>C. i</p> <p>D. i</p> <p>E. i i</p> |  <p>$X, \% = \frac{(V_{Na_2S_2O_3} - V_{Na_2S_2O_3}) \cdot K \cdot T \cdot 100}{m_H}$</p> <p>$T, / = \frac{I_2 \cdot s \cdot M}{1000}$</p> |
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| <p>« -2»</p> <p>1.</p> | <p>() :</p> |
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| <p>A. (III) B. C. 2,4- D. E.</p> |  $\text{C}_6\text{H}_5\text{OH} + \text{FeCl}_3 \longrightarrow \text{C}_6\text{H}_5\text{OFeCl}_2 + \text{HCl}$ |
| <p>2. A. B. C. D. E.</p> | $\text{R}-\overset{\text{O}}{\underset{\text{O}}{\text{C}}}-\text{OR}' \xrightarrow[\text{NaOH}]{\text{NH}_2\text{OH}} \text{R}-\overset{\text{O}}{\underset{\text{O}}{\text{C}}}-\text{NHOH} + \text{R}'-\text{OH}$ $3\text{R}-\overset{\text{O}}{\underset{\text{O}}{\text{C}}}-\text{NHOH} + \text{FeCl}_3 \longrightarrow \left(\text{R}-\overset{\text{O}}{\underset{\text{O}}{\text{C}}}-\text{NHO}\right)_3\text{Fe} + 3\text{HCl}$ <p style="text-align: right;">(Fe³⁺)</p> |
| <p>3. A. B. C. D. E.</p> | <p>(III)</p> $\text{RC}(=\text{O})\text{OR}' \xrightarrow{\text{NH}_2\text{OH}} \text{RC}(=\text{O})\text{NHOH} + \text{R}'\text{OH}$ $\text{RC}(=\text{O})\text{NHOH} \xrightarrow{\text{FeCl}_3} \left(\text{RC}(=\text{O})\text{NH-O}\right)_3\text{Fe} \cdot \left[\left(\text{RC}(=\text{O})\text{NH-O}\right)_2\text{Fe}\right]\text{Cl}$ |
| <p>4. A. B. C. D. E.</p> |  $\text{O}_2\text{N}-\text{C}_6\text{H}_4-\text{CH}(\text{OH})-\text{CH}(\text{NHC(=O)CHCl}_2)-\text{CH}_2\text{OH} + 4\text{NaOH} \longrightarrow \text{O}_2\text{N}-\text{C}_6\text{H}_4-\text{C}(\text{OH})_2-\text{CH}(\text{OH})-\text{CH}_2\text{OH} + \text{O}_2\text{N}-\text{C}_6\text{H}_4-\text{COONa} + \text{NH}_3 + \text{H}_2\text{O} + 2\text{NaCl}$ |
| <p>5. A. α- B. C. D. E.</p> | <p>a-</p>  $\text{Cyclic-CH}_2\text{-OH} \xrightarrow{[\text{O}]} \text{Cyclic-COOH}$ |

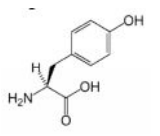
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| <p>1.</p> <p>A. i</p> <p>B. i</p> <p>C. i i</p> <p>D.</p> <p>E. i i</p> | $4\text{HCl} + \text{MnO}_2 \rightarrow \text{Cl}_2 \uparrow + \text{MnCl}_2 + 2\text{H}_2\text{O}$ |
| <p>2.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | $\text{Cl}_2 + 2\text{KI} \rightarrow \text{I}_2 + 2\text{KCl}$ <p>0,2 0,01</p> |
| <p>3.</p> <p>A.</p> <p>B.</p> <p>C. i i</p> <p>D.</p> <p>E.</p> | $\text{KI} + \text{AgNO}_3 \rightarrow \text{AgI} \downarrow + \text{KNO}_3$ $T = \frac{V_{\text{AgNO}_3} \cdot K_{\text{AgNO}_3} \cdot T \cdot V}{V_j} \quad T = \frac{c_{(\text{AgNO}_3)} \cdot S \cdot M_{(\text{KI})}}{1000} (\%)$ |
| <p>4.</p> <p>A. i</p> <p>B. i</p> <p>C. i i</p> <p>D. i</p> <p>E. i i</p> | $\text{K}^+ + \begin{array}{c} \text{COOH} \\ \\ \text{CH-OH} \\ \\ \text{CH-OH} \\ \\ \text{COOH} \end{array} \rightarrow \begin{array}{c} \text{COOK} \\ \\ \text{CH-OH} \\ \\ \text{CH-OH} \\ \\ \text{COOH} \end{array} + \text{H}^+$ |
| <p>5.</p> <p>A.</p> <p>D.</p> <p>E.</p> | $\text{I}_2 + 2\text{Na}_2\text{S}_2\text{O}_3 \rightarrow 2\text{NaI} + \text{Na}_2\text{S}_4\text{O}_6$ $\% = \frac{V_{\text{Na}_2\text{S}_2\text{O}_3} \cdot K \cdot T \cdot 100}{m} \quad T = \frac{c_{\text{Na}_2\text{S}_2\text{O}_3} \cdot s \cdot M_{(\text{I}_2)}}{1000} (\%)$ |
| <p>6.</p> <p>A.</p> <p>D.</p> |  |

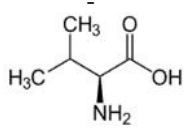
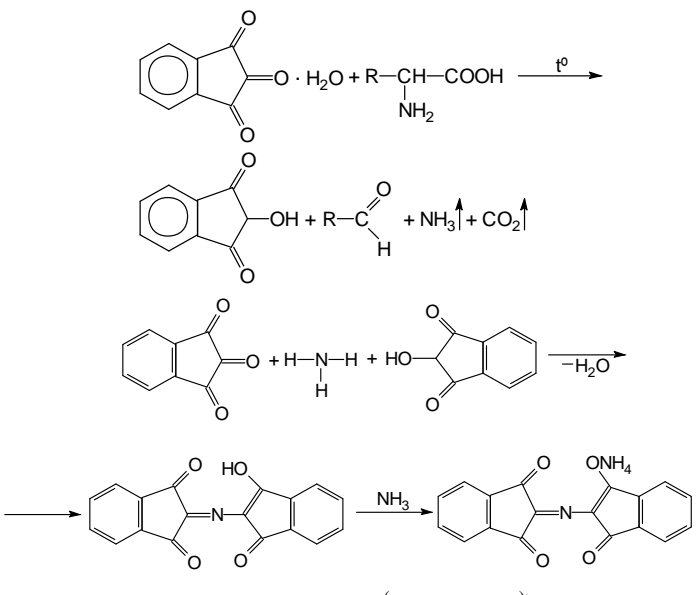
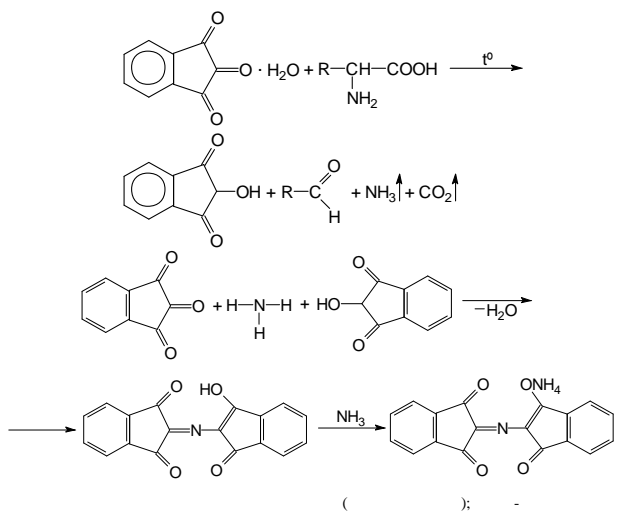
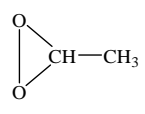
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| <p>E.</p> <p>7. i i , i i i i</p> <p>A. i B. i C. i D. i E.</p> | <p>, s = 1:</p> $\text{NaF} + \text{HClO}_4 \rightarrow \text{NaClO}_4 + \text{HF}$ $\% = \frac{(V - V) \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{(\text{HClO}_4)} \cdot s \cdot M}{1000} (\text{ } / \text{ })$ |
| <p>8. i i , i:</p> <p>A. i B. i C. i D. i E.</p> | <p>, s = 2,5:</p> $5\text{H}_2\text{O}_2 + 2\text{KMnO}_4 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{MnSO}_4 + \text{K}_2\text{SO}_4 + 8\text{H}_2\text{O} + 5\text{O}_2$ $\% = \frac{V_{\text{KMnO}_4} \cdot K \cdot T \cdot V \cdot 100}{V \cdot V}$ $T = \frac{c_{\text{KMnO}_4} \cdot s \cdot M_{(\text{H}_2\text{O}_2)}}{1000} (\text{ } / \text{ })$ |
| <p>9.</p> <p>50 5 0,3 ?</p> <p>A. B. C. D. E.</p> | $\text{MnO}_4^- + \text{C}_2\text{H}_5\text{OH} + 2\text{OH}^- \rightarrow \text{MnO}_4^{2-} + \text{H}_3\text{C}-\overset{\text{O}}{\parallel}{\text{C}}-\text{H} + 2\text{H}_2\text{O}$ $3\text{MnO}_4^{2-} + 2\text{H}_2\text{O} \rightarrow \text{MnO}_2\downarrow + 2\text{MnO}_4^- + 4\text{OH}^-$ |
| <p>10.</p> <p>3%.</p> <p>?</p> <p>A. B. C. D. E.</p> | <p>3%</p> $2\text{CrO}_4^{2-} + 2\text{H}^+ \rightarrow \text{Cr}_2\text{O}_7^{2-} + 2\text{H}_2\text{O}$ $\text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + \text{H}_2\text{O}_2 \rightarrow \text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + \text{H}_2\text{O}$ $\text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + 5\text{H}_2\text{O}_2 \rightarrow \text{HO}-\overset{\text{O}-\text{O}}{\parallel}{\text{Cr}}-\text{O}-\text{O}-\overset{\text{O}-\text{O}}{\parallel}{\text{Cr}}-\text{OH} + 5\text{H}_2\text{O}$ $\text{Cr}^{6+} \rightarrow \text{Cr}^{3+}$ |
| <p>11.</p> <p>A. B. C. D. E.</p> | <p>, s = 2,5:</p> $5\text{H}_2\text{O}_2 + 2\text{KMnO}_4 + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{MnSO}_4 + \text{K}_2\text{SO}_4 + 8\text{H}_2\text{O} + 5\text{O}_2$ $\% = \frac{V_{\text{KMnO}_4} \cdot K \cdot T \cdot V \cdot 100}{V \cdot V}$ $T = \frac{c_{\text{KMnO}_4} \cdot s \cdot M_{(\text{H}_2\text{O}_2)}}{1000} (\text{ } / \text{ })$ |
| <p>12. I i i i</p> <p>A. B. C. D. E.</p> | <p>Cr^{3+}</p> $2\text{CrO}_4^{2-} + 2\text{H}^+ \rightarrow \text{Cr}_2\text{O}_7^{2-} + 2\text{H}_2\text{O}$ $\text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + \text{H}_2\text{O}_2 \rightarrow \text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + \text{H}_2\text{O}$ |

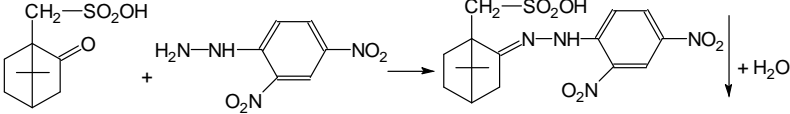
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| | $\text{HO}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{O}-\overset{\text{O}}{\parallel}{\text{Cr}}-\text{OH} + 5\text{H}_2\text{O}_2 \rightarrow \text{HO}-\overset{\text{O}-\text{O}}{\parallel}{\text{Cr}}-\text{O}-\overset{\text{O}-\text{O}}{\parallel}{\text{Cr}}-\text{OH} + 5\text{H}_2\text{O}$ $\text{Cr}^{6+} \rightarrow \text{Cr}^{3+}$ |
| <p>13.</p> <p>A. B. C. D. E.</p> | <p>Na₂S₂O₃ + 2HCl → 2NaCl + SO₂↑ + S↓ + H₂O</p> <p>5SO₂ + 2KIO₃ → I₂ + 4SO₃ + K₂SO₄</p> |
| <p>14.</p> <p>A. B. C. D. E.</p> | $\text{KBrO}_3 + 5\text{KBr} + 6\text{HCl} \rightarrow 3\text{Br}_2 + 6\text{KCl} + 3\text{H}_2\text{O}$  <p>s=3 100</p> $\frac{2,353 \cdot (a - b)}{m}$ <p>— , 0.0167</p> <p>b — , 0.0167</p> <p>m —</p> $2,353 = \frac{T \cdot V_M \cdot 100}{V} = \frac{0,009428 \cdot 25 \cdot 100}{10}$ <p>100</p> <p>40</p> |
| <p>15.</p> <p>A. B. C. D. E.</p> | $\text{MgCO}_3 + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$ |
| <p>16.</p> <p>A. D. E.</p> | <p>s = 1:</p> $\text{Ca}^{2+} + \text{H}_2 \text{Ind} \rightarrow [\text{Ca Ind}] + 2\text{H}^+$  |

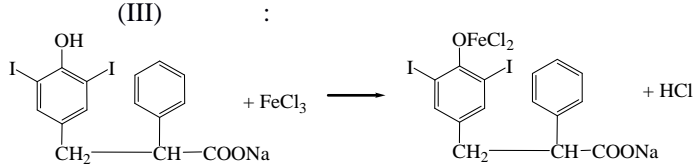
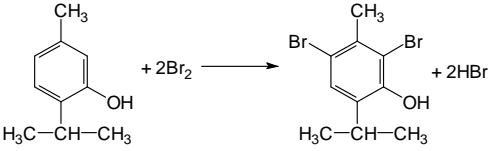
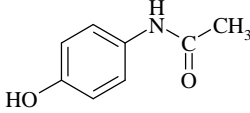
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| | $[\text{Ca Ind}] + \begin{array}{l} \text{H}_2\text{C}-\text{N} \begin{array}{l} \text{CH}_2\text{COONa} \\ \text{CH}_2\text{COOH} \end{array} \\ \\ \text{H}_2\text{C}-\text{N} \begin{array}{l} \text{CH}_2\text{COOH} \\ \text{CH}_2\text{COONa} \end{array} \end{array} \longrightarrow \left[\begin{array}{l} \text{H}_2\text{C}-\text{N} \begin{array}{l} \text{CH}_2\text{COONa} \\ \text{CH}_2\text{COO} \end{array} \\ \\ \text{H}_2\text{C}-\text{N} \begin{array}{l} \text{CH}_2\text{COO} \\ \text{CH}_2\text{COONa} \end{array} \end{array} \text{Ca} \right] + 2\text{H}^+$ $\% = \frac{V}{m} \cdot K \cdot T \cdot 100 \quad T = \frac{c_{()} \cdot s \cdot M_{()}}{1000} (\ / \ /)$ |
| <p>17.</p> <p>A. .</p> <p>D. -</p> <p>E.</p> | <p>(II)</p> <p>:</p> $\text{Hg}^{2+} + 2\text{NaOH} \rightarrow \text{HgO} \downarrow + 2\text{Na}^+ + \text{H}_2\text{O}$ <p>?</p> |
| <p>18.</p> <p>A. .</p> <p>D. -</p> <p>E.</p> | <p>:</p> $\text{Ag}^+ + \text{HCl} \rightarrow \text{AgCl} \downarrow + \text{H}^+$ |
| <p>19. i i i i (II)</p> <p>i i i i</p> <p>A. i i?</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E. i</p> | <p>,</p> <p>:</p> <p>, s = 1:</p> $3\text{FeSO}_4 + 3(\text{NH}_4)_2 \text{e}(\text{NO}_3)_6 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{Fe}(\text{NO}_3)_3 + 3\text{Ce}(\text{NO}_3)_3 + 6\text{NH}_4\text{NO}_3$ $\left[\left(\begin{array}{c} \text{C}_5\text{H}_4\text{N}_2 \\ \text{C}_5\text{H}_4\text{N}_2 \\ \text{C}_5\text{H}_4\text{N}_2 \end{array} \right) \text{Fe} \right]^{2+} + \text{Ce}^{4+} \longrightarrow \left[\left(\begin{array}{c} \text{C}_5\text{H}_4\text{N}_2 \\ \text{C}_5\text{H}_4\text{N}_2 \\ \text{C}_5\text{H}_4\text{N}_2 \end{array} \right) \text{Fe} \right]^{3+} + \text{Ce}^{3+}$ $\% = \frac{V_{(\text{NH}_4)_2\text{Ce}(\text{NO}_3)_6} \cdot K \cdot T \cdot 100}{m} \quad T = \frac{c_{(\text{NH}_4)_2\text{Ce}(\text{NO}_3)_6} \cdot s \cdot M_{()}}{1000} (\ / \ /)$ |
| <p>20.</p> <p>A. .</p> <p>D. .</p> <p>E.</p> | <p>812 17.10.2012 «</p> <p>()</p> <p>»</p> <p>).</p> <p>(</p> |

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| <p>1.</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">:</p> $\text{C}_2\text{H}_5\text{OH} + 4\text{I}_2 + 6\text{NaOH} \rightarrow \text{CHI}_3\downarrow + 5\text{NaI} + \text{HCOONa} + 5\text{H}_2\text{O}$ |
| <p>2.</p> <p>A. i i (II)</p> <p>B. i (II)</p> <p>C. (II)</p> <p>D. i (II) i</p> <p>E. i i</p> | <p style="text-align: center;">()</p> $\text{CuSO}_4 + 2\text{NaOH} \rightarrow \text{Cu(OH)}_2 + \text{Na}_2\text{SO}_4$ $2 \begin{array}{c} \text{H}_2\text{C}-\text{OH} \\ \\ \text{HC}-\text{OH} \\ \\ \text{H}_2\text{C}-\text{OH} \end{array} + \text{Cu(OH)}_2 \rightarrow \begin{array}{c} \text{H}_2\text{C}-\text{O} \\ \\ \text{HC}-\text{O} \\ \\ \text{H}_2\text{C}-\text{OH} \end{array} \begin{array}{c} \text{H} \\ \diagup \\ \text{O}-\text{CH}_2 \\ \\ \text{O}-\text{CH} \\ \\ \text{HO}-\text{CH}_2 \end{array} \text{Cu} + 2\text{H}_2\text{O}$ |
| <p>3.</p> <p>A. ?</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">()</p> |
| <p>4.</p> <p>A. i?</p> <p>D. (III)</p> <p>E.</p> | <p style="text-align: center;">:</p> $\text{K}^+ + \begin{array}{c} \text{COOH} \\ \\ \text{CHOH} \\ \\ \text{CHOH} \\ \\ \text{COOH} \end{array} \rightarrow \begin{array}{c} \text{COOK} \\ \\ \text{CHOH} \\ \\ \text{CHOH} \\ \\ \text{COOH} \end{array} + \text{H}^+$ |
| <p>5.</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">()</p> <p style="text-align: center;">; s = 1:</p> $\text{CH}_3\text{COOK} + \text{HClO}_4 \xrightarrow{\text{C}_3} \text{CH}_3\text{COOH} + \text{KClO}_4$ $\% = \frac{(V - V') \cdot K \cdot T \cdot 100}{m} \quad T = \frac{c_{(\text{HClO}_4)} \cdot s \cdot M}{1000} \quad (/)$ |
| <p>6.</p> <p>A. ?</p> | |

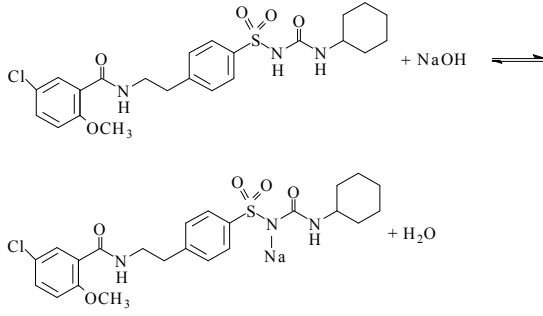
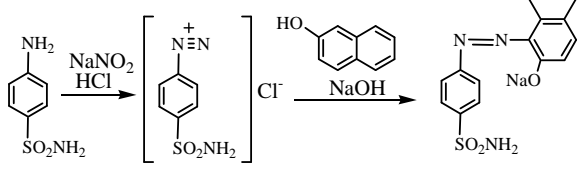
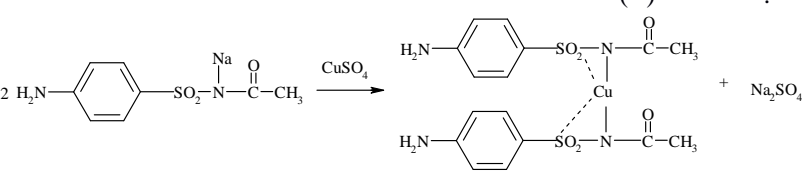
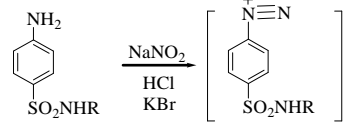
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| <p>D. E.</p> | |
| <p>7.</p> <p>A. B. C. D. E.</p> | <p style="text-align: center;"> $\text{HOOC-CH}_2\text{-CH}_2\text{-CH(NH}_2\text{)-COOH} + \text{C}_6\text{H}_4(\text{CO})_2 \cdot \text{H}_2\text{O} \xrightarrow{t^\circ} \text{HOOC-CH}_2\text{-CH}_2\text{-CHO} + \text{C}_6\text{H}_4(\text{CO})_2\text{-OH} + \text{NH}_3 \uparrow + \text{CO}_2 \uparrow$ </p> <p style="text-align: center;"> $\text{C}_6\text{H}_4(\text{CO})_2 + \text{H-NH-H} + \text{HO-C}_6\text{H}_4(\text{CO})_2 \xrightarrow{-\text{H}_2\text{O}}$ </p> <p style="text-align: center;"> $\rightarrow \text{C}_6\text{H}_4(\text{CO})_2\text{-N=C}_6\text{H}_4(\text{CO})_2 \xrightarrow{\text{NH}_3} \text{C}_6\text{H}_4(\text{CO})_2\text{-N=C}_6\text{H}_4(\text{CO})_2\text{-ONH}_4$ </p> <p style="text-align: center;">(); -</p> |
| <p>8.</p> <p>A. D. E.</p> | <p style="text-align: center;">  </p> <p style="text-align: center;">2- -3-(4-)</p> <p style="text-align: center;"> $\text{C}_6\text{H}_4(\text{CO})_2 \cdot \text{H}_2\text{O} + \text{R-CH(NH}_2\text{)-COOH} \xrightarrow{t^\circ}$ </p> <p style="text-align: center;"> $\text{C}_6\text{H}_4(\text{CO})_2\text{-OH} + \text{R-C(=O)-H} + \text{NH}_3 \uparrow + \text{CO}_2 \uparrow$ </p> <p style="text-align: center;"> $\text{C}_6\text{H}_4(\text{CO})_2 + \text{H-NH-H} + \text{HO-C}_6\text{H}_4(\text{CO})_2 \xrightarrow{-\text{H}_2\text{O}}$ </p> <p style="text-align: center;"> $\rightarrow \text{C}_6\text{H}_4(\text{CO})_2\text{-N=C}_6\text{H}_4(\text{CO})_2 \xrightarrow{\text{NH}_3} \text{C}_6\text{H}_4(\text{CO})_2\text{-N=C}_6\text{H}_4(\text{CO})_2\text{-ONH}_4$ </p> <p style="text-align: center;">(); -</p> |

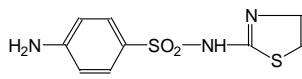
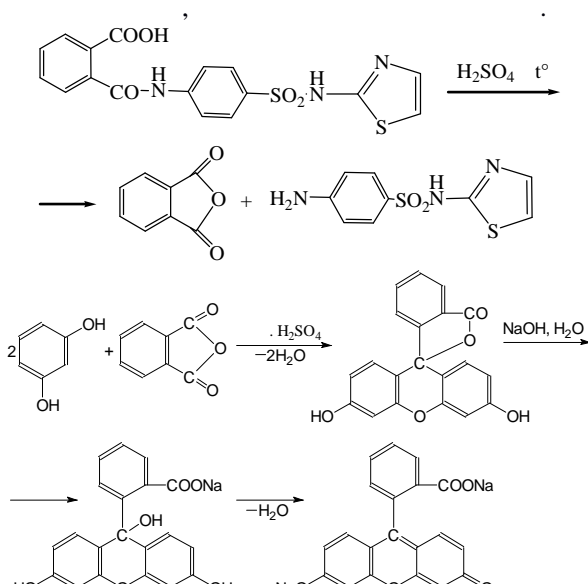
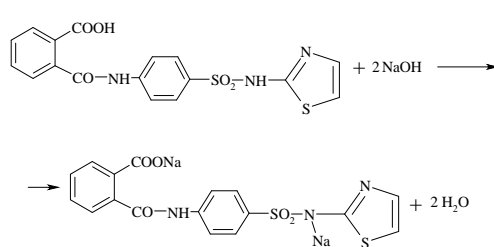
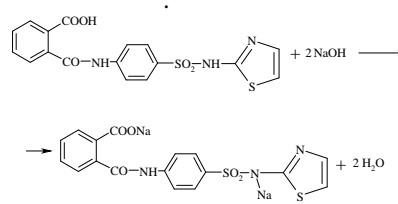
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| <p>9.</p> <p>A. .2,4- D. 2,4- E.</p> |  <p>2- -3-</p>  <p>() ; -</p> |
| <p>10.</p> <p>A. B. D. E.</p> |  <p>() ; -</p> |
| <p>11. i i i - i i i i i i: A. i B. i D. i E. i</p> | $C_2H_5-O-C_2H_5 \xrightarrow{O_2} C_2H_5-O-O-C_2H_5$  |

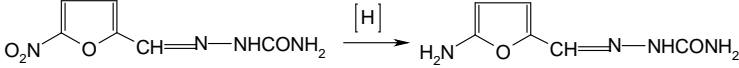
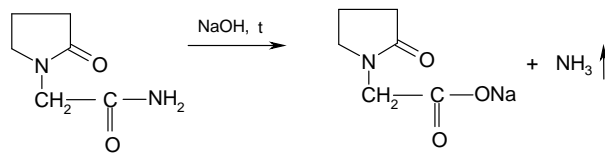
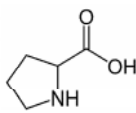
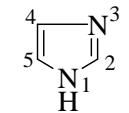
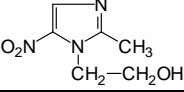
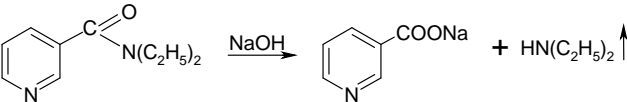
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| <p>12. i i i - i i i :</p> <p>A. i</p> <p>B. i</p> <p>C. i</p> <p>D. i i</p> <p>E. i i</p> | <p style="text-align: center;">:</p> $\text{C}_2\text{H}_5\text{-O-O-C}_2\text{H}_5 + 2\text{KI} + \text{H}_2\text{O} \longrightarrow \text{I}_2 + \text{C}_2\text{H}_5\text{-O-C}_2\text{H}_5 + 2\text{KOH}$ |
| <p>13. -</p> <p>().</p> <p>?</p> <p>A. 0,1</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">,</p> $\begin{array}{c} \text{H}_5\text{C}_6 \\ \\ \text{CH-O-CH}_2\text{CH}_2\text{N(CH}_3)_2 \cdot \text{HCl} \\ \\ \text{H}_5\text{C}_6 \end{array} \xrightarrow{\text{H}_2\text{SO}_4}$ $\longrightarrow \left(\begin{array}{c} \text{H}_5\text{C}_6 \\ \\ \text{CH-O}^+\text{-CH}_2\text{CH}_2\text{-N}^+\text{(CH}_3)_2 \\ \quad \\ \text{H} \quad \text{H} \end{array} \right) \text{SO}_4^{2-} + \text{HCl}$ |
| <p>14. 180°C</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">2 15° . 15°</p> <p>20° . 180°C :</p> $4\text{C}_3\text{H}_5(\text{ONO}_2)_3 \longrightarrow 6\text{N}_2\uparrow + 12\text{CO}_2\uparrow + \text{O}_2\uparrow + 10\text{H}_2\text{O}$ <p style="text-align: right;">!</p> |
| <p>15. i i i i</p> <p>i i i . ?</p> <p>A. i</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">-</p> $\begin{array}{c} \text{NH}_2 \\ \\ \text{O=C} \\ \\ \text{NH}_2 \end{array} \longrightarrow \text{NH=C=O} + \text{NH}_3\uparrow$ $\text{NH=C=O} + \begin{array}{c} \text{NH}_2 \\ \\ \text{O=C} \\ \\ \text{NH}_2 \end{array} \longrightarrow \begin{array}{c} \text{H}_2\text{N-C-NH-C-NH}_2 \\ \quad \\ \text{O} \quad \text{O} \end{array}$ $\begin{array}{c} \text{O=C-NH-C=O} \\ \quad \\ \text{NH}_2 \quad \text{NH}_2 \end{array} \rightleftharpoons \begin{array}{c} \text{O=C-NH-C-OH} \\ \quad \\ \text{NH}_2 \quad \text{NH} \end{array} \xrightarrow{\text{Cu}^{2+}; \text{KOH}} \begin{array}{c} \text{O=C-NH-C-OK} \\ \quad \\ \text{H}_2\text{N} \quad \text{N} \\ \diagdown \quad / \\ \text{Cu/2} \end{array}$ |
| <p>16. i - i i i i</p> <p>i i i i . 2,4- i i</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D. i</p> <p>E.</p> | <p style="text-align: center;">2,4-</p>  <p style="text-align: center;">:</p> |

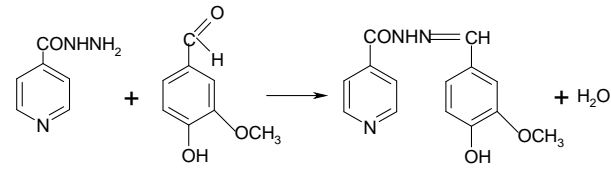
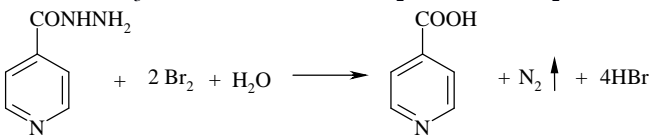
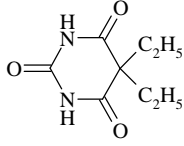
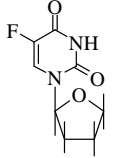
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| <p style="text-align: center;">« -2»</p> | |
| <p>1.</p> <p style="text-align: center;">(III)</p> <p>A. : (III)</p> <p>B. .</p> <p>C. .</p> <p>D. .</p> <p>E. .</p> |  <p style="text-align: center;">+ FeCl₃ → + HCl</p> |
| <p>2.</p> <p>A. .</p> <p>B. .</p> <p>C. .</p> <p>D. .</p> <p>E. .</p> | <p>0,1</p> <p style="text-align: center;">, s = 1,5:</p> $\text{KBrO}_3 + 5\text{KBr} + 6\text{HCl} \rightarrow 3\text{Br}_2 + 6\text{KCl} + 3\text{H}_2\text{O}$  <p style="text-align: center;">+ 2HBr</p> $\% = \frac{V_{\text{KBrO}_3} \cdot K \cdot T \cdot 100}{m}$ $T = \frac{C_{(\text{KBrO}_3)} \cdot S \cdot M_{()}}{1000} (/)$ |
| <p>3.</p> <p>A. ():</p> <p>B. .</p> <p>C. .</p> <p>D. .</p> <p>E. .</p> | |
| <p>4.</p> <p>(),</p> <p>A. N-(4-)</p> <p>B. -</p> <p>C. -</p> <p>D. 2-()</p> <p>E. 2-[[2,6-()]-]]</p> | <p style="text-align: right;">N-(4-</p>  |
| <p>5.</p> <p>A. .</p> <p>B. .</p> <p>C. .</p> <p>D. .</p> <p>E. .</p> | <p>0,1</p> <p>00 (00,) ; s=1:</p> |

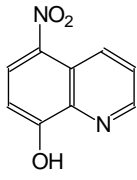
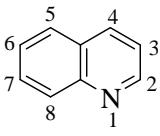
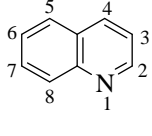
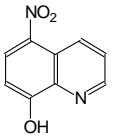
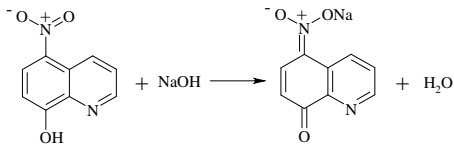
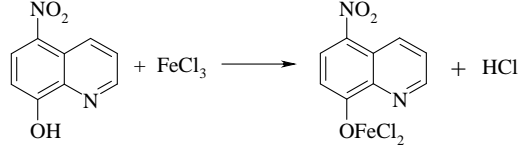
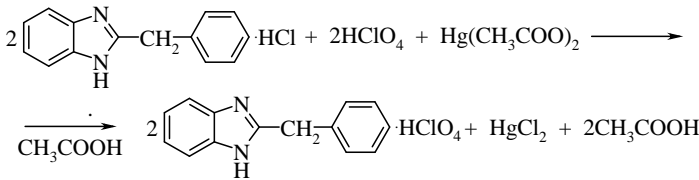
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|---|---|
| | <p style="text-align: center;"> $5\text{NaNO}_2 + 2\text{KIO}_3 + 2\text{HCl} \rightarrow \text{I}_2 + 5\text{NaNO}_3 + 2\text{KCl} + \text{H}_2\text{O}$ </p> <p style="text-align: center;"> $\% = \frac{(V - V_0) \cdot K \cdot T \cdot 100}{m}$ </p> <p style="text-align: center;"> $T = \frac{c_{(\text{NaNO}_2)} \cdot s \cdot M}{1000}$ </p> |
| <p>6.</p> <p>A.</p> <p>D.</p> <p>E.</p> | |
| <p>7.</p> <p>A. FeCl_3</p> <p>B. $\text{K}_2[\text{gI}_4]$</p> <p>C. $[\text{NH}_4]_2\text{C}_2\text{O}_4$</p> <p>D. KMnO_4</p> <p>E. K_2CrO_4</p> | <p style="text-align: center;">(III)</p> |
| <p>8.</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;"> $\% = \frac{V_{\text{HClO}_4} \cdot K \cdot T \cdot 100}{m}$ </p> <p style="text-align: center;"> $T = \frac{c_{(\text{HClO}_4)} \cdot s \cdot M}{1000}$ </p> |
| <p>9.</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p style="text-align: center;">; s = 1.</p> <p style="text-align: center;"> $2\text{KI} + \text{Br}_2 \rightarrow \text{I}_2 + 2\text{KBr}$ </p> <p style="text-align: center;"> $\text{I}_2 + 2\text{Na}_2\text{S}_2\text{O}_3 \rightarrow 2\text{NaI} + \text{Na}_2\text{S}_4\text{O}_6$ </p> <p style="text-align: center;"> $\% = \frac{(V - V_0) \cdot K \cdot T \cdot 100}{m}$ </p> <p style="text-align: center;"> $T = \frac{c_{(\text{KBrO}_3)} \cdot s \cdot M}{1000}$ </p> |
| <p>10. salicylas)</p> <p style="text-align: right;">(Phenylii</p> | |

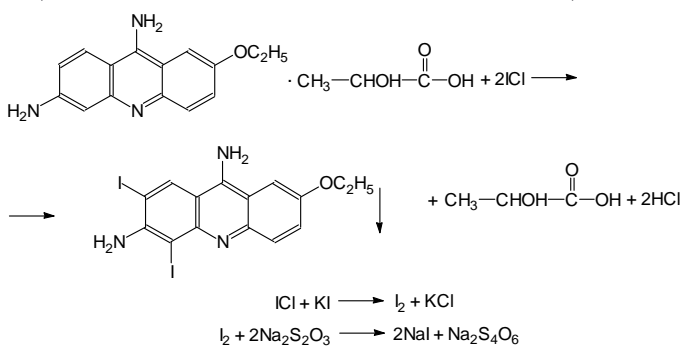
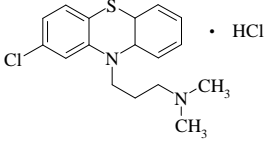
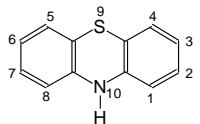
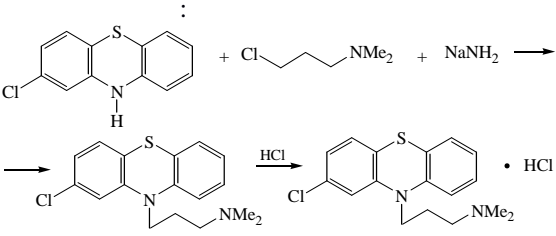
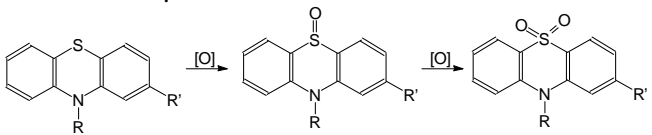
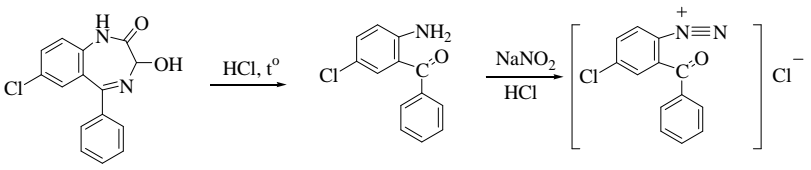
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| <p>16.</p> <p>A. B. C. D. E.</p> | <p>, s = 1:</p>  $T = \frac{C_{(NaOH)} \cdot s \cdot M_{(\quad)}}{1000} (/) \quad \% = \frac{V_{NaOH} \cdot K_{NaOH} \cdot T \cdot 100}{m}$ |
| <p>17.</p> <p>A. B. C. D. E.</p> | |
| <p>18.</p> <p>A. B. C. D. E.</p> |  |
| <p>19.</p> <p>A. B. C. D. E.</p> |  |
| <p>20.</p> <p>A. B. C. D. E.</p> |  $T = \frac{C_{(NaNO_2)} \cdot s \cdot M_{(\quad)}}{1000} (/) \quad \% = \frac{V_{NaNO_2} \cdot K_{NaNO_2} \cdot T \cdot 100}{m}$ $\% = \frac{(V_{NaNO_2} - V_{NaNO_2}) \cdot K_{NaNO_2} \cdot T \cdot 100}{m}$ |

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| <p>21.</p> <p>A B. C. D. E.</p> | <p>()</p> <p>:</p>  <p>$H_2S + Pb(CH_3COO)_2 \rightarrow PbS + 2CH_3COOH$</p> |
| <p>22.</p> <p>A. D. E.</p> |  |
| <p>23.</p> <p>A. D. E.</p> | <p>() ; s =</p> <p>1/2 ():</p>  $\% = \frac{(V_{NaOH} - V_{NaOH}) \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{(NaOH)} \cdot s \cdot M_{()}}{1000} (/)$ |
| <p>24.</p> <p>A. B. C. D. E.</p> | <p>()</p>  $\% = \frac{(V_{NaOH} - V_{NaOH}) \cdot K \cdot T \cdot 100}{m} \quad T = \frac{c_{(NaOH)} \cdot s \cdot M_{()}}{1000} (/)$ |

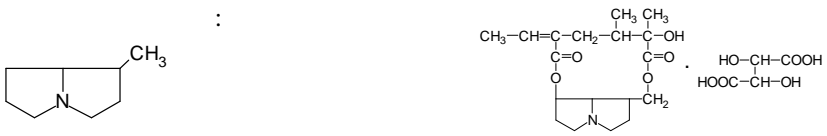
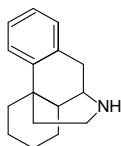
| « -2» | |
|--|--|
| <p>1. 5- : A. : D. E.</p> | <p>() 5- : </p> |
| <p>2. () - ? A. B. F l₂ C. F l₃ D. E. gN</p> | <p>:', : </p> |
| <p>3. - (), ? A. B. C. D. E.</p> | <p>: () : </p> |
| <p>4. : 2-(2- -5- -1 - -1-) ? A. B. C. D. E.</p> | <p>-, 1,3:  -1 - -1-) 2-(2- -5- </p> |
| <p>5. i i i i i i i i i i i i i i i , i i . i ' i i i : A. i i B. i i C. i D. E. i i</p> | <p>:', : </p> |
| <p>6. i - i i i i i i i i i i</p> | <p>(), :</p> |

| | |
|---|--|
| <p>A. i i</p> <p>B. i i i</p> <p>C. i i</p> <p>D. i i i</p> <p>E. i i</p> |  |
| <p>7.</p> <p>A.</p> <p>D.</p> <p>E.</p> | <p>$\text{KBrO}_3 + 5\text{KBr} + 6\text{HCl} \rightarrow 3\text{Br}_2 + 6\text{KCl} + 3\text{H}_2\text{O}$</p>  <p>$\% = \frac{V_{\text{KBrO}_3} \cdot K \cdot T \cdot 100}{m_H}$</p> <p>$T = \frac{c_{(\text{KBrO}_3)} \cdot s \cdot M}{1000} (\text{ / })$</p> |
| <p>8.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | <p>5,5-</p>  |
| <p>9.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | <p>(NaOH).</p> |
| <p>10. (Phlorafur)</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> |  <p>$2\text{F}^- + \text{CaCl}_2 \rightarrow \text{CaF}_2 + 2\text{Cl}^-$</p> |

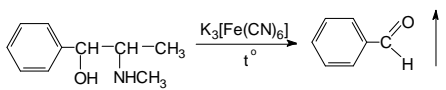
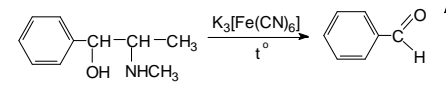
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| <p>11.</p> <p>A. ?</p> <p>B. (</p> <p>C.) -1,4-</p> <p>D.</p> <p>E. 4-</p> |  <p>5- -8-</p>  |
| <p>12.</p> <p>A. : 5- -8-</p> <p>D.</p> <p>E.</p> |  <p>8- -8-</p>  |
| <p>13.</p> <p>A. - :</p> <p>D.</p> <p>E. -</p> |  |
| <p>14.</p> <p>A. i (III)</p> <p>B. i i</p> <p>C. i i (II)</p> <p>D. i</p> <p>E.</p> | <p>(III)</p>  |
| <p>15.</p> <p>A. i?</p> <p>B. i i</p> <p>C.</p> <p>D. i i</p> <p>E. i i</p> | <p>(,), - , HCl,</p> <p>() , s=1.</p>  $T = \frac{C_{HClO_4} \cdot s \cdot M(\quad)}{m} (\% /) \quad \% = \frac{(V_{\dots} - V_{\dots}) \cdot K \cdot T \cdot 100}{m}$ |

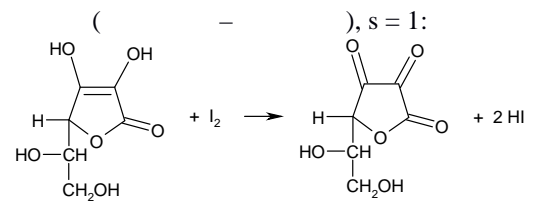
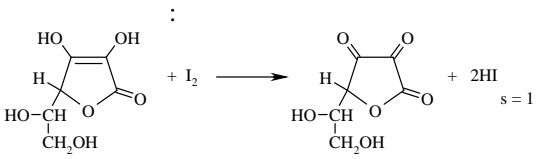
| | |
|--|--|
| <p>16.</p> <p>i i i</p> <p>A. i i i</p> <p>B. i i i</p> <p>C. i 00 i</p> <p>D. i</p> <p>E.</p> | <p>· CH₃-CHOH-C(=O)-OH + 2HCl →</p> <p>→</p> <p>+ CH₃-CHOH-C(=O)-OH + 2HCl</p> <p>ICl + KI → I₂ + KCl</p> <p>I₂ + 2Na₂S₂O₃ → 2NaI + Na₂S₄O₆</p> <p>$\% = \frac{(V_{Na_2S_2O_3} - V_{Na_2S_2O_3}) \cdot K \cdot T \cdot 100}{m}$</p> <p>$T = \frac{C_{ICl} \cdot s \cdot M(\quad)}{m} (\quad / \quad)$</p> <p>, s=1/2:</p>  |
| <p>17.</p> <p>i i i</p> <p>A. i</p> <p>B. i</p> <p>C. I</p> <p>D. i</p> <p>E.</p> |  <p>(2- -10-(3-))</p>  |
| <p>18.</p> <p>A. 2-</p> <p>B. 5-</p> <p>C. 4-</p> <p>D. 3-</p> <p>E. 6-</p> | <p>2-</p>  |
| <p>19.</p> <p>A. i</p> <p>B. i</p> <p>C. i i</p> <p>D. i (III)</p> <p>E. i</p> |  |
| <p>20.</p> <p>A.</p> <p>B.</p> <p>C.</p> |  |

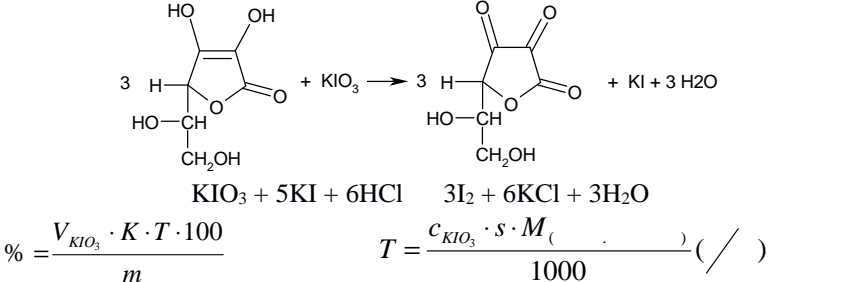
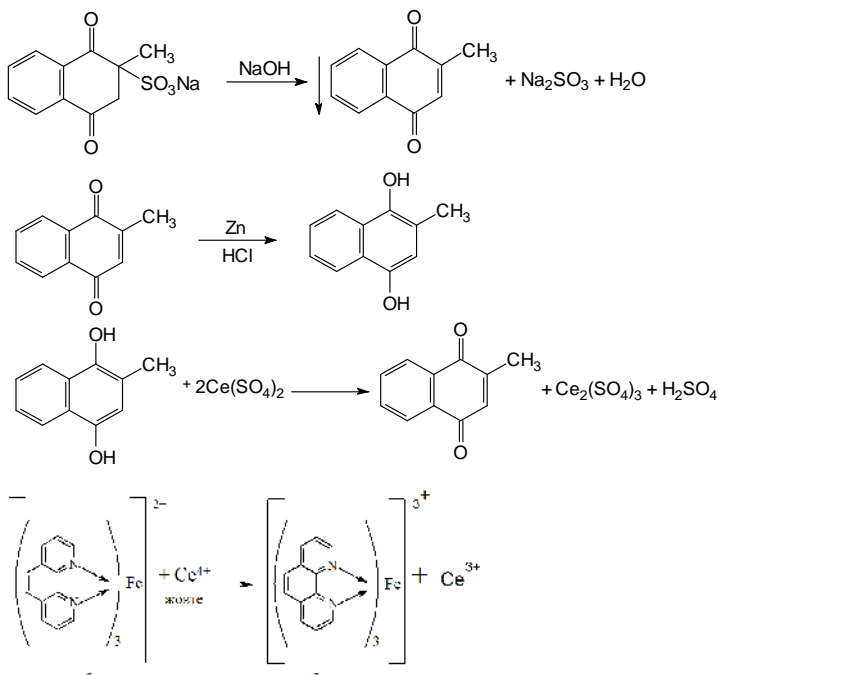
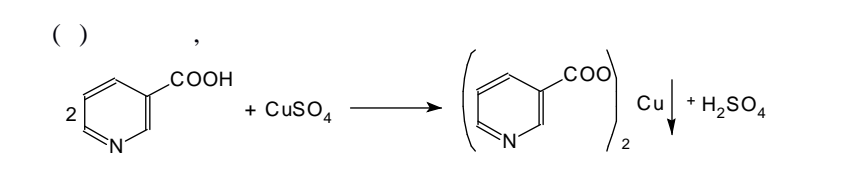
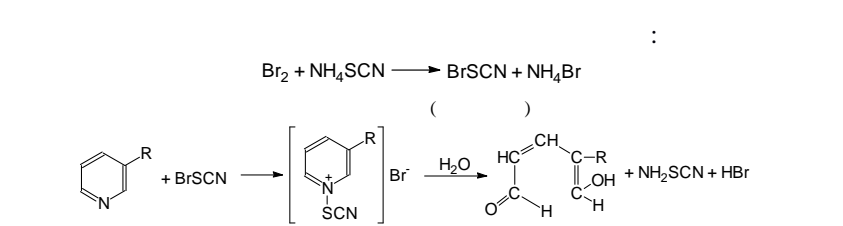
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| D. | gN | |
| E. | | |
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| | () | . |
| 230 | 350 . | |
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| | : | |
| A. | | |
| B. | | |
| C. | | |
| D. | | |
| E. | - | |

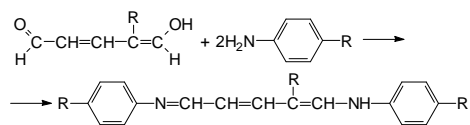
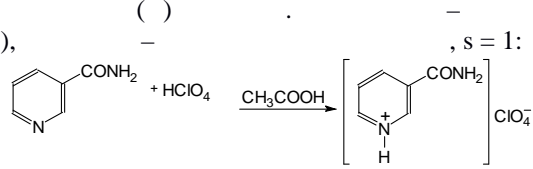
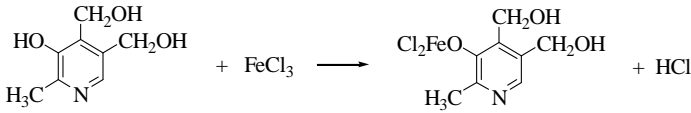
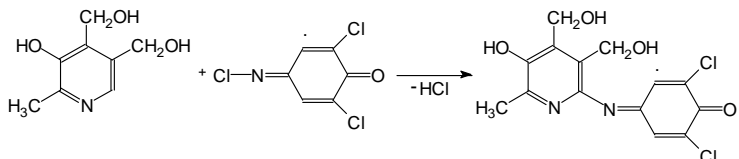
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|----|------------------------|--|
| | « -2» | |
| 1. | | ([iL4]). |
| A. | : | |
| B. | | |
| C. | | |
| D. | | |
| E. | | |
| 2. | i i i i i i : | (s=1): |
| A. | | |
| B. | i | |
| C. | i | |
| D. | i i | |
| E. | i | |
| | | $\left[\begin{array}{c} \text{CH}_2\text{OH} \\ \\ \text{N}^+\text{-CH}_3 \\ \\ \text{H} \end{array} \right] \text{O-C-CH-C}_6\text{H}_5 \left. \vphantom{\left[\begin{array}{c} \text{CH}_2\text{OH} \\ \\ \text{N}^+\text{-CH}_3 \\ \\ \text{H} \end{array} \right]} \right]_2 \text{SO}_4^{2-} + \text{HClO}_4 \xrightarrow{\text{CH}_3\text{COOH}}$ $\rightarrow \left[\begin{array}{c} \text{CH}_2\text{OH} \\ \\ \text{N}^+\text{-CH}_3 \\ \\ \text{H} \end{array} \right] \text{O-C-CH-C}_6\text{H}_5 \text{ClO}_4^- + \left[\begin{array}{c} \text{CH}_2\text{OH} \\ \\ \text{N}^+\text{-CH}_3 \\ \\ \text{H} \end{array} \right] \text{O-C-CH-C}_6\text{H}_5 \text{HSO}_4^-$ $\% = \frac{V_{\text{HClO}_4} \cdot K \cdot T \cdot 100 \cdot 100}{m \cdot (100 - \%)} \quad T = \frac{C_{(\text{HClO}_4)} \cdot s \cdot M_{()}}{1000} (/)$ |
| 3. | i i i i, i i i : | 1- |
| A. | i i i | |
| B. | i | |
| C. | i | |
| D. | i i | |
| E. | i i i | |
| | |  |
| 4. | , | |
| A. | : | |
| B. | | |
| C. | | |
| D. | | |
| E. | | |
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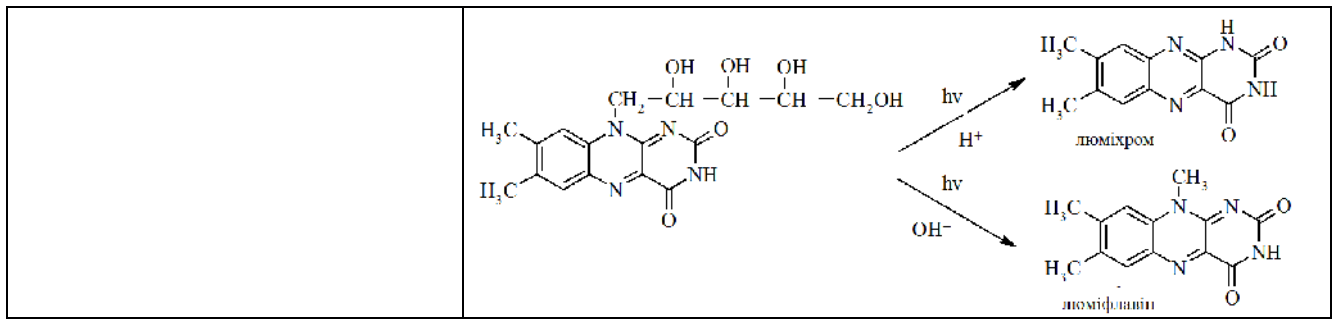
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|---|---|
| | <p>7,8- -4,5α- -3- -17- -6α-</p> |
| <p>5.</p> <p>-3,7- -1 - 1,3,7- -2,6-</p> <p>A. : . . D. : E.</p> | <p>() () :</p> <p>1,3,7- -3,7- -1 - -2,6- :</p> |
| <p>6.</p> <p>- ?</p> <p>A. ? . . D. - E.</p> | <p>).</p> <p>0,1</p> <p>(S = 1);</p> $\% = \frac{V_{HClO_4} \cdot K \cdot 100 \cdot 100}{m \cdot (100 - \%)} \quad T = \frac{c_{(HClO_4)} \cdot s \cdot M}{1000} \quad (/)$ |
| <p>7.</p> <p>- ?</p> <p>A. ? . . D. ? E.</p> | |
| <p>8.</p> <p>i i - i i</p> <p>i i i i i</p> <p>i i . i</p> <p>i i ?</p> <p>A. (II) . . D. : E.</p> | <p>() , -</p> <p>:</p> $2 \begin{array}{c} \text{CH}_2 - \text{NH}_2 \\ \\ \text{CH}_2 - \text{NH}_2 \end{array} + \text{CuSO}_4 \longrightarrow \left[\text{Cu} \left(\begin{array}{c} \text{CH}_2 - \text{NH}_2 \\ \\ \text{CH}_2 - \text{NH}_2 \end{array} \right)_2 \right] \text{SO}_4$ |
| <p>9.</p> <p>i i</p> <p>i i i i</p> <p>i</p> <p>?</p> | <p>:</p> |

| | |
|--|--|
| <p>A. i B. i C. i D. i i E.</p> |  |
| <p>10.</p> <p>A. B. C. D. E.</p> | <p>():</p>  |

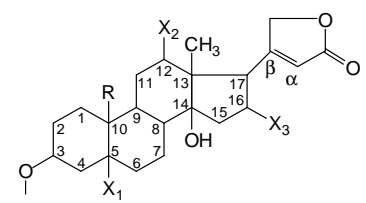
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|---|---|
| <p style="text-align: center;">« -2 »</p> | |
| <p>1.</p> <p>A. B. C. D. E.</p> | <p style="text-align: center;">:</p> $\text{H}_2\text{C}_2\text{O}_4 + \text{CaCl}_2 \rightarrow \text{CaC}_2\text{O}_4 + 2\text{HCl}$ |
| <p>2.</p> <p>A. B. C. D. E.</p> | <p style="text-align: center;">,</p>  <p style="text-align: center;">), s = 1:</p> $\% = \frac{V_{I_2} \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{(I_2)} \cdot s \cdot M(\text{...})}{1000} (\text{...})$ |
| <p>3.</p> <p>A. B. C. D. E.</p> | <p style="text-align: center;">:</p>  <p style="text-align: center;">), s = 1:</p> $\% = \frac{V_{I_2} \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{(I_2)} \cdot s \cdot M(\text{...})}{1000} (\text{...})$ |
| <p>4.</p> <p>A. B. C.</p> | <p style="text-align: center;">,</p> <p style="text-align: center;">- , s = 3.</p> |

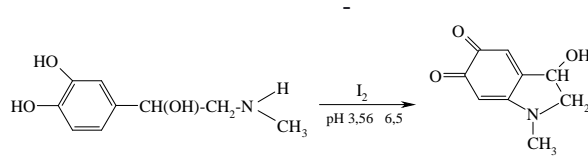
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| <p>D. E.</p> |  |
| <p>5. A. . D. E.</p> | <p style="text-align: right;">; s = 1/2.</p> <p style="text-align: right;">2- -1,4- , 2- -1,4- - (IV)</p>  <p style="text-align: center;"> $\% = \frac{(V - V_{\text{жёлте}}) \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{\text{Ce}(\text{SO}_4)_2} \cdot s \cdot M}{1000} \quad (/)$ </p> |
| <p>6. A. . D. E.</p> |  |
| <p>7. A. . D. E.</p> |  |

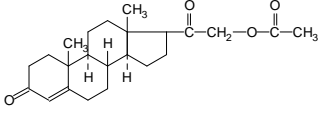
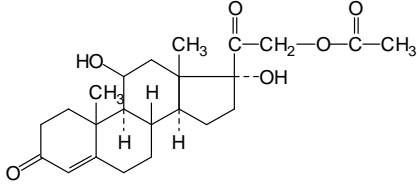
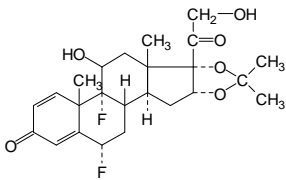
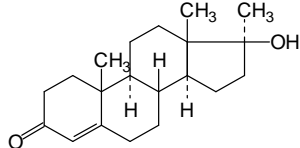
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| <p>8.</p> <p>A. i i i</p> <p>B. i i i</p> <p>C. i i i (II)</p> <p>D. i i i</p> <p>E. i i i</p> | <p>(), () , s = 1:</p>  $\% = \frac{V_{HClO_4} \cdot K \cdot T \cdot 100}{m}$ $T = \frac{c_{(HClO_4)} \cdot s \cdot M}{1000} (\text{ } / \text{ })$ |
| <p>9.</p> <p>A. (III)</p> <p>D.</p> <p>E.</p> | <p>(III)</p>  |
| <p>10.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> | <p>2,6-</p>  |
| <p>11.</p> <p>A. i i i i?</p> <p>B. i i i</p> <p>C. i i i</p> <p>D. i i i i</p> <p>E. i i i</p> | |
| <p>12.</p> <p>A. i i</p> <p>B. i i</p> <p>C. i i</p> <p>D. i i</p> <p>E. i i</p> | |

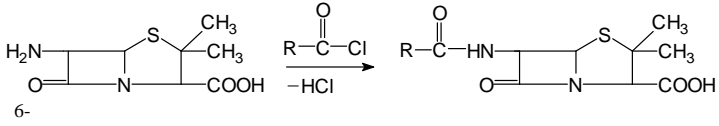
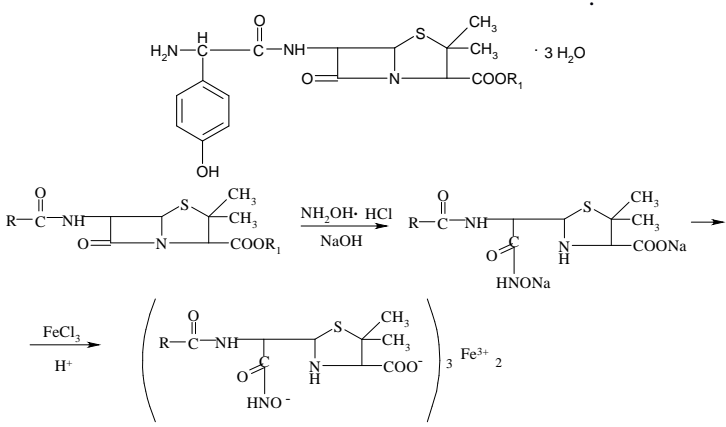
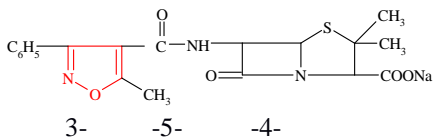
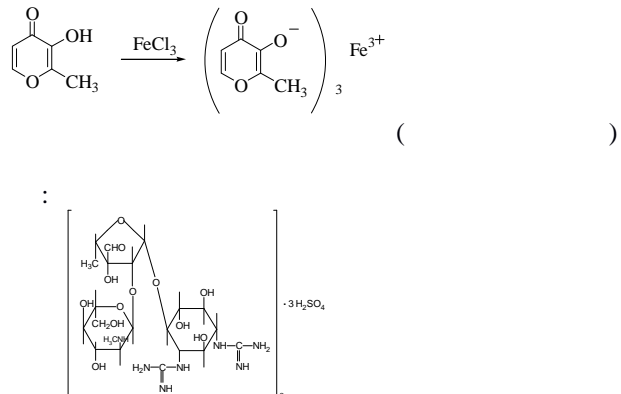
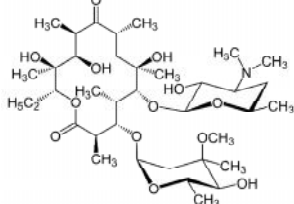


| « -2» | |
|---|---|
| <p>1.</p> <p>A. .</p> <p>D. .</p> <p>E. .</p> | <p style="text-align: center;">():</p> $ \begin{array}{c} \text{C=O} \\ \\ \text{H-C-OH} \\ \\ \text{HO-C-H} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{CH}_2\text{OH} \end{array} + 2 \begin{array}{c} \text{NaOOC-CH-CH-COO}^- \\ \quad \\ \text{HO} \quad \text{OH} \\ \diagup \quad \diagdown \\ \text{Cu}^{2+} \\ \diagdown \quad \diagup \\ \text{HO} \quad \text{OH} \\ \quad \\ \text{OOC-CH-CH-COOK} \end{array} + 3 \text{NaOH} + 2\text{KOH} \longrightarrow $ $ \longrightarrow \begin{array}{c} \text{COONa} \\ \\ \text{H-C-OH} \\ \\ \text{HO-C-H} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{CH}_2\text{OH} \end{array} + 4 \begin{array}{c} \text{COONa} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{COOK} \end{array} + 2 \text{CuOH} + 2\text{H}_2\text{O} $ $ 2\text{CuOH} \xrightarrow{t} \text{Cu}_2\text{O} \downarrow + \text{H}_2\text{O} $ |
| <p>2.</p> <p>A. .</p> <p>D. .</p> <p>E. .</p> | <p style="text-align: center;">():</p> $ \begin{array}{c} \text{C=O} \\ \\ \text{H-C-OH} \\ \\ \text{HO-C-H} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{CH}_2\text{OH} \end{array} + 2 \begin{array}{c} \text{NaOOC-CH-CH-COO}^- \\ \quad \\ \text{HO} \quad \text{OH} \\ \diagup \quad \diagdown \\ \text{Cu}^{2+} \\ \diagdown \quad \diagup \\ \text{HO} \quad \text{OH} \\ \quad \\ \text{OOC-CH-CH-COOK} \end{array} + 3 \text{NaOH} + 2\text{KOH} \longrightarrow $ $ \longrightarrow \begin{array}{c} \text{COONa} \\ \\ \text{H-C-OH} \\ \\ \text{HO-C-H} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{CH}_2\text{OH} \end{array} + 4 \begin{array}{c} \text{COONa} \\ \\ \text{H-C-OH} \\ \\ \text{H-C-OH} \\ \\ \text{COOK} \end{array} + 2 \text{CuOH} + 2\text{H}_2\text{O} $ $ 2\text{CuOH} \xrightarrow{t} \text{Cu}_2\text{O} \downarrow + \text{H}_2\text{O} $ |
| <p>3.</p> <p>A. .</p> <p>D. .</p> <p>E. .</p> | <p style="text-align: center;">D- , ≈63% β-D-(+)- , ≈36% α-D-</p> <p style="text-align: center;">(+)-</p> |

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| <p>4.</p> <p>A. B. C. D. E.</p> | <p>()</p> |
| <p>5.</p> <p>A. B. C. D. E.</p> | <p>()</p> |
| <p>6.</p> <p>A. B. C. D. E.</p> | <p>17</p> <p>()</p>  |
| <p>7.</p> <p>(G),</p> <p>A. B. C. D. E.</p> | <p>(G)</p> |

| | |
|---------------------------------|--|
| <p>1.</p> <p>0,05</p> <p>A.</p> | <p>« -2»</p> <p>0,05</p> <p>(6,5):</p>  |
|---------------------------------|--|

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|---|---|
| <p>D. E.</p> | |
| <p>2. i i i - i i i i i i i i ? A. D. E.</p> | <p>(Desoxycorticosteroni acetat)</p>  <p>-4- -21- -3,20-21-</p> |
| <p>3. A. B. C. D. E.</p> | <p>4-5).</p> |
| <p>4. -3,2- 11 ,17- -4- -21- A. B. C. D. E.</p> | <p>-4- -21- 11 ,17- -3,2-</p>  |
| <p>5. A. D. E.</p> | <p>6α, 9α- -16α- -16,17-</p>  |
| <p>6. i i i i A. B. i C. i i D. i E.</p> | <p>17α- -4- -17β- -3</p>  |

| « -2» | |
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| <p>1.</p> <p>A. 6-</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E. 7-</p> | <p>6-</p>  |
| <p>2.</p> <p>(III)</p> <p>A.</p> <p>D.</p> <p>E.</p> |  <p>(III)</p> |
| <p>3.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> |  <p>3- -5- -4-</p> |
| <p>4.</p> <p>A.</p> <p>B.</p> <p>C.</p> <p>D.</p> <p>E.</p> |  <p>()</p> |
| <p>5.</p> <p>A. β-</p> <p>D.</p> <p>E.</p> |  <p>12-17</p> |

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« -2 »

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